Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Ionic Compounds**

|  |  |  |  |
| --- | --- | --- | --- |
|  | **Ionic Formula** | **Name of Ionic Compound** | **Balanced Ion Pairs** |
| 1. | Al2O3 | Aluminum oxide | 2Al3+ + 3O2- |
| 2. | Ba3N2 |  |  |
| 3. | Cs2O |  |  |
| 4. |  | Indium fluoride |  |
| 5. |  |  | Mg2+ + O2- |
| 6. |  | Calcium oxide |  |
| 7. | Rb3N |  |  |
| 8. |  |  | Ca2+ + 2Cl- |
| 9. |  | Sodium nitride |  |
| 10. | Ca3N2 |  |  |
| 11. |  |  | 2Li+ + O2- |
| 12. |  | Magnesium chloride |  |
| 13. | SrSe |  |  |
| 14. |  |  | 3 Na+ + P3- |
| 15. | Cs2S |  |  |
| 16. |  |  | Be2+ + 2Cl- |
| 17. |  | Potassium oxide |  |
| 18. | Al2S3 |  |  |

19. Explain the difference between a cation and an anion. Explain how each is formed. Provide an example of each.

20. In your own words, explain how to balance ion pairs. (Pretend you are teaching someone who has no background in atoms, ions, charges, losing/gaining electrons)